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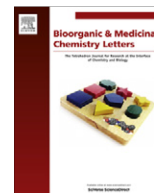

Synthesis and structure–activity relationship of piperidine-derived non-urea soluble epoxide hydrolase inhibitors

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ABSTRACT

A series of potent amide non-urea inhibitors of soluble epoxide hydrolase (sEH) is disclosed. The inhibition of soluble epoxide hydrolase leads to elevated levels of epoxyeicosatrienoic acids (EETs), and thus inhibitors of sEH represent one of a novel approach to the development of vasodilatory and anti-inflammatory drugs. Structure–activities studies guided optimization of a lead compound, identified through high-throughput screening, gave rise to sub-nanomolar inhibitors of human sEH with stability in human liver microsomal assay suitable for preclinical development.

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Soluble epoxide hydrolase (sEH) is a bifunctional homodimeric enzyme with hydrolase and phosphatase activity detected in various species ranging from plants to mammals.[1] In humans it is mostly located in liver, kidney, intestinal and vascular tissues.[2] The sEH enzyme is selective for aliphatic epoxides of fatty acids, and one of the most important substrates is epoxyeicosatrienoic acid (EET).[3] EETs are one of the metabolic derivatives of arachidonic acid.[4] The epoxygenase CYP2 enzymes catalyse the epoxydation of olefin bonds of arachidonic acid generating EETs.[5] EETs exhibit vasodilatory effects in various arteries[6,7] and possess anti-inflammatory properties.[8] sEH mediates the addition of water to EETs, converting them to the corresponding diols—dihydroxyeicosatrienoic acids (DHETs), which show abolished or diminished biological activity. The inhibition of this enzyme, therefore may represent a promising therapeutic strategy since it would lead to elevated levels of EETs, which could then have beneficial therapeutic effects on blood pressure and inflammation.[9]

The most explored sEH inhibitors to date are urea-based compounds and numerous structure–activity relationship (SAR) studies led to discovery of several potent urea-based sEH inhibitors with IC_{50}s in the lower nanomolar range.[10,11] However, the urea-based inhibitors often suffer from poor solubility and stability, which hinders their pharmacological use in vivo.[12,13]

Amides, carbamates and thioureas have been evaluated as alternative pharmacophores in an attempt to improve physical properties of urea-based inhibitors.[10,11] In this report we disclose our efforts towards designing novel non-urea, amide-based inhibitors of sEH.

Previously we identified through high-throughput screening (HTS)[14] several hits with a low micromolar to low nanomolar potency. A majority of these compounds were urea-based inhibitors. However, several were non-ureas and among them the most potent was the derivative of isonipecotic acid, with substantial activity of 20 nM. Our initial SAR studies[15] led to discovery of compound **1** with an IC_{50} of 7.9 nM (Fig. 1). Based on its potent inhibition of sEH activity, we continued our evaluation of the right-hand side (2,4,6-trimethylphenyl ring) of **1**, and this follow-up SAR study revealed compound **2**, with an IC_{50} of 1.6 nM.[16]

Herein we report modifications of the left-hand side of non-urea inhibitor **2**. In particular, we were focused on the replacement

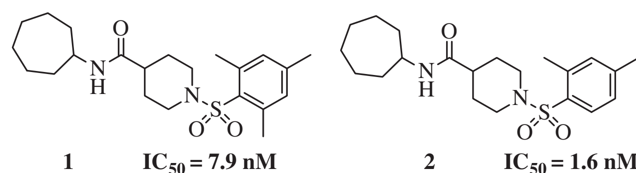

Figure 1. Chemical structures of non-urea inhibitors.

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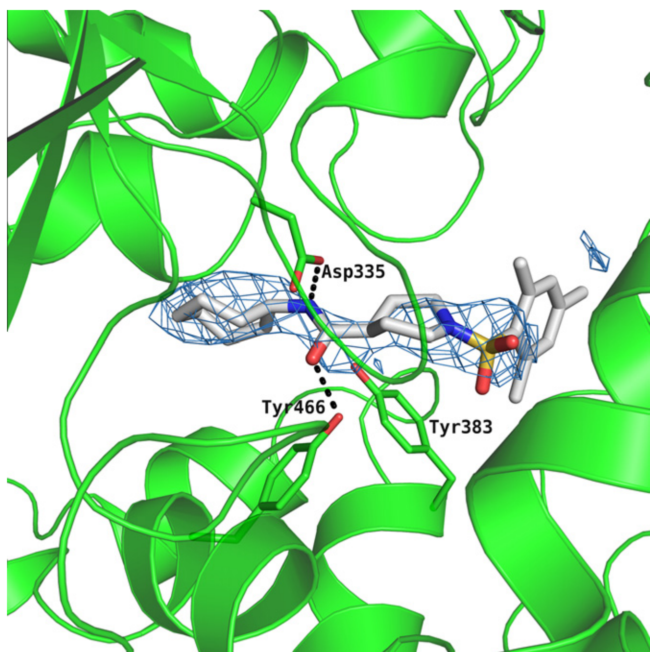


Figure 2. The active site of the hydrolase domain of human sEH complexed with **1**. The F_o–F_c electron density map is contoured at 3σ. Hydrogen bonds are indicated by dashes. The structure reveals that the 2,4,6-trimethylphenyl group is directed into the hydrophobic environment, whereas the cycloheptyl moiety is exposed toward the solvent. Coordinates of the complex deposited as 4HAI with the Protein Data Bank. The image was produced using PYMOL.

of the cycloheptyl moiety of the **2**, following the SAR reported by Rose et al.[17], and guided by docking model of the previously reported protein-inhibitor complex and X-ray cocrystal structure of human sEH and **1**.

The X-ray crystallographic structure of human sEH and an inhibitor 4-(3-cyclohexylureido)-carboxylic acid complex (PDB code: 1ZD3) revealed the catalytic pocket and key structural features required to inhibit sEH enzyme. Two tyrosine (Tyr383 and Tyr466) and one aspartic acid (Asp335) residues, located in the hydrolase catalytic pocket of sEH, are involved in degradation of EET—tyrosine residues act as hydrogen bond donors to promote the epoxide ring opening by Asp335.[18,19] The urea group of 4-(3-cyclohexylureido)-carboxylic acid fits in the hydrolase catalytic pocket and the carbonyl oxygen of the urea moiety is engaged in a hydrogen bond interaction with Tyr381 and Tyr466 while the N–H acts as a hydrogen bond donor to Asp335 The co-crystal structure of **1** and human sEH was determined (See Supplementary data) to guide the design of more potent inhibitors. Our close examination of the structure revealed that the amide moiety of **1** is positioned in the same fashion as the urea group in urea-based inhibitors, where the amide moiety, instead of urea group, is involved in hydrogen bonding with tyrosine and aspartic acid residues in the catalytic pocket of sEH. We noticed that the 2,4,6-trimethylphenyl ring on the right-hand side of this inhibitor occupies the smaller of the two lipophilic pockets in the sEH active site, while the cycloheptyl moiety on the left-hand side is directed towards the large deep pocket that opens toward solvent, which allows access for further structural modifications (Fig. 2). We therefore hypothesized that the left-hand segment in **1** could be optimized to further improve the potency of this group of non-urea sEH inhibitors.

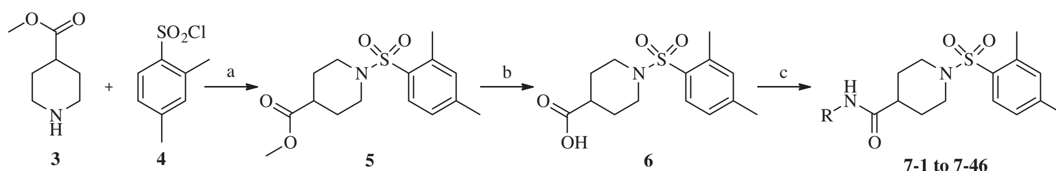
Scheme 1 outlines the synthetic route used to form the left-hand side library of non-urea amide sEH inhibitors. Sulfonamide **5** was prepared from two commercially available starting materials, methyl isonipecotate **3** and 2,4-dimethylbenzenesulfonyl chloride **4**. Saponification of this methyl ester with LiOH afforded acid **6**. EDC peptide coupling reactions of **6** with various commercially available amines provided analogs **7–1** to **7–46**.

A sensitive fluorescent based assay was employed to determine IC_{50} values of these sEH inhibitors. In short, cyano(2-methoxynaphthalen-6-yl)methyl trans-(3-phenyloxyran-2-yl) methyl carbonate (CMNPC) was used as the fluorescent substrate. Human sEH (1 nM) was incubated with the inhibitor for 5 min in pH 7.0 Bis–Tris/HCl buffer (25 mM) containing 0.1 mg/mL of BSA at 30 °C prior to substrate introduction ([*S*] = 5 μM). Activity was determined by monitoring the appearance of 6-methoxy-2-naphthaldehyde over 10 min by fluorescence detection with an excitation wavelength of 330 nm and an emission wavelength of 465 nm. Reported IC_{50} values are the average of the three replicates with at least two datum points above and at least two below the IC_{50}.[20] In Table 1 are summarized the biological results.

Previous studies on urea-based inhibitors containing piperidine moiety have shown that the hydrophobic cycloalkyl groups on the left-hand side of the molecules are positively correlated with inhibitory potency.[17] Among this set of analogs we identified several inhibitors possessing improved or similar potency comparing to the lead compound **2**, specifically compound **7–10** showed an IC_{50} of 0.4 nM, the most potent amide non-urea sEH inhibitor reported to date. Replacement of cycloalkyl ring with a more compact phenyl ring (compound **7–12**), resulted in 15-fold drop in potency against the human sEH. However, introduction of the phenyl ring allowed us access to electronically and sterically diverse structures, and as well attachment of various polar groups, which could in turn improve physical and pharmacokinetic properties. Placement of fluorine or bromine in the *ortho* position did not significantly change the potency of the non-urea inhibitors (**7–13** and **7–15**), while chlorine and methyl group decreased the potency for 10 and 30-fold, respectively (**7–14** and **7–16**).

Polar hydroxyl group in *ortho* position showed clear negative effect on potency in non-urea based compounds (**7–17**). Although the *para* substitution is generally tolerated, placement of polar substituents resulted in less potent inhibitors.

Placement of methoxy group in *para* position (compound **7-18**) did not significantly changed the potency comparing to compound **7-12**, while introduction of hydroxyl group in the same position (compound **7-19** can be observed as a metabolite of **7-18**) led to two fold decreased potency presumably because of unfavorable electron character and increased polarity. Similar results, but with more drastically change in potency can be observed based on results for methyl ester compound **7-21** and its corresponding carboxylic acid compound **7-22**. We synthesized the 4-trifluoromethoxyphenyl analog **7-23**, since it has been reported previously that



Scheme 1. Reagents and condition: (a) Et_3N, CH_2Cl_2, rt 24 h; 89% (b) LiOH, THF/H_2O, rt, 24 h; 91% (c) $R-NH_2$, EDC, DMAP, CH_2CL, rt, 24 h; 65–92%.

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