Contents lists available at ScienceDirect





Fluid Phase Equilibria

PC-SAFT based group contribution method for binary interaction parameters of polymer/solvent systems



Felix T. Peters, Franziska S. Laube, Gabriele Sadowski*

Laboratory of Thermodynamics, Department of Biochemical and Chemical Engineering, TU Dortmund, Emil-Figge-Str. 70, 44227 Dortmund, Germany

ARTICLE INFO

Article history: Received 16 November 2012 Received in revised form 4 March 2013 Accepted 28 May 2013 Available online 29 June 2013

Keywords: PC-SAFT Group contribution method Binary interaction parameters Polymers Phase equilibrium

$A \hspace{0.1in} B \hspace{0.1in} S \hspace{0.1in} T \hspace{0.1in} R \hspace{0.1in} A \hspace{0.1in} C \hspace{0.1in} T$

Based on the PC-SAFT equation of state, a group-contribution method is developed for the determination of binary interaction parameters (k_{ij}) between polymers and solvents (GCM- k_{ij}). Applying simple geometric combination rules, the k_{ij} values can be calculated from binary interaction parameters between polymer groups and solvent groups (k_{ps}). 80 different k_{ps} values allow the calculation of a k_{ij} for polymer/solvent systems with polymers consisting of $-CH_3$, CH_2 , $|CH_{2,HDPE}$, $|(CH)-, |CC_i)$, $|CH_{a,r}, |C-_{ar}$ and -(C=O)-O-groups and solvents containing $-CH_3$, $|CH_2, \rangle$ ($CH_2, |_{CPC}$), $|(CH)-_{Cyclo}$, $=CH_{2,alkene}, =(CH)-_{alkene}, |CH_{Ar}, |C-_{Ar}, -O-, |C=O$ and -(C=O)-O-groups. Using earlier-published group-contributions for polymer parameters (Peters et al. [2,3]) in combination with the proposed GCM- k_{ij} for polymer/solvent systems allows for predicting liquid-liquid phase equilibria, vapor-liquid phase equilibria and excess enthalpies of binary polymer/solvent and copolymer/solvent as well as ternary polymer/solvent 1/solvent 2 systems in good agreement with experimental data.

© 2013 Elsevier B.V. All rights reserved.

1. Introduction

Quick but reliable predictions of thermo-physical properties of polymers and copolymers as well as of binary (co-)polymer/solvent and ternary (co-)polymer/solvent 1/solvent 2 systems are highly requested in polymer industry. At this, modeling provides a useful tool which replaces cost-intensive and time-consuming measurements of experimental data.

In the literature various equations of state (EOS) have been developed and checked for polymer systems. Among these the Perturbed Chain Statistical Associating Fluid Theory (PC-SAFT) by Gross and Sadowski [1] offers the advantage to account for the chain-like structure of polymers. Calculation of requested systems requires PC-SAFT model parameters for each considered compound. For polymers they are usually fitted to experimental density data of the pure polymer as well as binary phase-equilibrium data with at least one solvent. Often, hardly any experimental data and even less already-fitted polymer parameters can be found in the literature. Therefore, a group-contribution method for polymer parameters (GCM-P) was developed in two previous works [2,3]. It was shown to be applicable to the modeling of polymers which contain the following groups $-CH_3$, CH_2 , (CH)-, C(, -(C=0)-O-, CH_{Ar} , $C_{Ar}-$, -O-, C=O, -OH, and/or Si(.

Modeling binary or higher phase-equilibrium data requires the use of a binary interaction parameter (k_{ij}) . For most polymer systems, this correction factor for polymer/solvent interactions differs from zero and has to be fitted to experimental data. This makes pure predictions of polymer/solvent phase equilibria not reliable even when the pure-component PC-SAFT parameters of involved substances are already known.

Due to that, in this work we propose a group-contribution method for the prediction of polymer/solvent k_{ij} values (GCM- k_{ij}). Using GCM-P polymer parameters, published solvent parameters and the k_{ij} determined by this GCM- k_{ij} , binary (co-)polymer/solvent and polymer/solvent 1/solvent 2 systems were studied in this work. Liquid–liquid phase equilibria (LLE), vapor–liquid phase equilibria (VLE) as well as excess enthalpies (h^E) were considered for modeling calculations.

2. PC-SAFT

PC-SAFT calculates the residual Helmholtz energy (a^{res}) of a system. This property allows the calculation of any other thermo-physical properties including phase equilibria using basic thermodynamic relations. Depending on the chemical behavior of the involved components, PC-SAFT considers various contributions to a^{res} (Eq. (1)). Since both, polymers and solvents, were considered as being nonpolar within this work and only non-associating components were taken into account, only two contributions to the Helmholtz energy were applied here. These describe the repulsive interactions by means of a hard-chain reference fluid (a^{hc}) and

^{*} Corresponding author. Tel.: +49 231 7552635; fax: +49 231 7552572. *E-mail address:* Gabriele.Sadowski@bci.tu-dortmund.de (G. Sadowski).

^{0378-3812/\$ -} see front matter © 2013 Elsevier B.V. All rights reserved. http://dx.doi.org/10.1016/j.fluid.2013.05.033

the dispersive van-der-Waals interactions between the molecules (a^{disp}) .

$$a^{\rm res} = a - a^{\rm id} = a^{\rm hc} + a^{\rm disp} \tag{1}$$

For PC-SAFT modeling, three parameters are required for characterizing each component in the system. A physically reasonable description is obtained using the segment number m^{seg} and the temperature-independent segment diameter σ to define the size of the molecule. An adequate representation of its dispersive, attractive interaction behavior is characterized by the dispersion-energy parameter u/k_B .

For a more-detailed description of the distinct contributions to the residual Helmholtz energy and the implementation of the pure-component parameters in the model equations the reader is referred to the original PC-SAFT publication of Gross and Sadowski [1].

An extension of the PC-SAFT to be used for copolymer systems was developed by Gross et al. [4] and was utilized within this work when copolymer systems were considered. A detailed description of the theory is renounced here and can be found in the original work.

For the modeling of multi-component systems the purecomponent parameters of the involved substances are combined according to the van-der-Waals one-fluid theory [5]. In this way the mean segment number \bar{m}_{seg} is determined as a weighted average of the segment numbers of components *i* in the requested system.

$$\bar{m}^{\text{seg}} = \sum_{i} x_i m_i^{\text{seg}} \tag{2}$$

The combination of the segment diameters and the dispersion energies of a pair of unlike segments is done applying the Berthelot–Lorentz combining rules [6]. Here the segment diameters of component i and j are averaged arithmetically whereas the dispersion energy of the mixture is determined by a geometric mean.

$$\sigma_{ij} = \frac{1}{2}(\sigma_i + \sigma_j) \tag{3}$$

$$\frac{u_{ij}}{k_B} = \sqrt{\frac{u_i}{k_B} - \frac{u_j}{k_B}(1 - k_{ij})}$$
(4)

The parameter k_{ij} is called binary interaction parameter and is used as a correction factor for the dispersive interaction of a binary mixture consisting of components *i* and *j*. It is usually fitted to binary experimental data of the corresponding mixture.

3. Group-contribution method for polymer parameters: GCM-P

Inspired by the simple arithmetic and geometric form of the Berthelot–Lorentz combining rules, a group-contribution method for polymer parameters (GCM-P) was proposed in [2,3]. According to this, polymer parameters are calculated from group contributions using Eqs. (5)-(7).

$$\frac{u_i}{k_B} = \frac{\sum_{k=1}^{n_{\text{groups},i}} n_{k,i}}{\sqrt{\prod_{k=1}^{n_{\text{groups},i}}} \left(\frac{u_k}{k_B}\right)^{n_{k,i}}}$$
(5)

$$\sigma_{i} = \frac{\sum_{k=1}^{n_{\text{groups},i}} n_{k,i} \sigma_{k}}{\sum_{k=1}^{n_{\text{groups},i}} n_{k,i}}$$
(6)

$$m_i^{\text{seg}} = \sum_{k=1}^{n_{\text{groups},i}} n_{k,i} m_k^{\text{seg}}$$
(7)

Here $n_{k,i}$ is the number of group k in polymer i. The sums and products are evaluated for all n_{groups} different groups in polymer i.

Parameterization of poly(olefins), poly(acrylates), poly(methacrylates) and further aromatic, oxygen- and siliconbased polymers was achieved using polymer groups $-CH_3$, CH_2 , $(CH)-, C(, -(C=0)-O-, CH_{Ar}, C_{Ar}-, -O-, C=0, -OH and/or)Si(.$ A detailed description of the parameterization procedure for polymers as well as the GCM-P parameters are listed in corresponding tables of the previous publications [2,3].

However, since the GCM-P only provides pure-component polymer parameters, k_{ij} values were still fitted to binary experimental data.

4. Group contribution method for binary interaction parameters: GCM-*k*_{ii}

To obtain a fully-predictive method for polymer/solvent systems, in this work a second group-contribution method (GCM- k_{ij}) is proposed to estimate the binary interaction parameters k_{ij} . In fact, this is inspired by the calculation of the dispersion-energy parameter in a mixture based on the Berthelot–Lorentz combining rule (Eq. (8)). Using interaction parameters k_{ps} between polymer groups p and solvent groups *s* of a polymer/solvent system allows for the determination of the polymer/solvent binary parameter k_{ij} (derivation is given in Appendix A):

$$(1 - k_{ij}) = \begin{bmatrix} \sum_{p=1}^{n_{groups,i}} n_{p,i} * \sum_{s=1}^{n_{groups,j}} n_{s,j} \end{bmatrix} \sqrt{\prod_{p=1}^{n_{groups,j}} \prod_{s=1}^{n_{groups,j}} (1 - k_{ps})^{n_{p,i} * n_{s,j}}}$$
(8)

Here, binary interaction parameters k_{ps} between the different polymer and solvent groups are combined applying a geometric mixing rule.

For applying Eq. (8), the polymers were divided into same polymer groups as defined earlier [2,3]. For the GCM- k_{ij} determination, the solvent molecule is also divided into solvent-specific groups. However, the solvent groups were only used for the estimation of the k_{ij} whereas the pure-component solvent parameters were not determined from group contributions. Due to large solvent-parameter tables in the literature and extensively experimental solvent data for parameters.

The $k_{\rm ps}$ values for all investigated combinations of polymer groups and solvent groups within this work were determined by minimizing the deviations between the k_{ij} values obtained by Eq. (8) and the k_{ij} values that were obtained as a best manual fit to binary experimental polymer/solvent data.

Since poly(olefins) in solutions of linear alkanes and 1-alkenes were assumed to be most ideal with respect to their dispersive interactions and due to the best availability of experimental data for these systems, they were chosen as starting systems for the specification of the group contributions. Doing this, k_{ps} values were determined for $-CH_3$, CH_2 , polymer groups with –CH₃,)CH₂,)(CH)–, =CH_{2,alkene}, =(CH)–_{alkene} solvent groups. Maintaining this set of binary group contributions, poly(acrylates) and poly(methacrylates) dissolved in linear alkanes and 1-alkenes were studied next and group contributions were specified for solvent interactions with the -(C=O)-O- polymer group. For this purpose, (acrylate)- and (methacrylate)-based polymers were considered with up to four aliphatic polymer groups in the side-chain of the monomer unit, i.e. PBA and PBMA. In a subsequent step, poly(styrene) and its derivates were considered. At the same time, the variety of considered solvents was extended toward cycloalkanes. Thus, new polymer/solvent group contributions were identified for the polymer groups >CHAr and C_{Ar} , with solvent groups $CH_{2,Cvclo}$ and $(CH)_{Cvclo}$. Finally, new GCM-kii contributions were determined for solutions of Download English Version:

https://daneshyari.com/en/article/203265

Download Persian Version:

https://daneshyari.com/article/203265

Daneshyari.com