



Review

Visible-light activation of TiO₂ photocatalysts: Advances in theory and experiments



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ARTICLE INFO

Article history:

Received 25 April 2015

Received in revised form 17 August 2015

Accepted 24 August 2015

Available online 28 August 2015

Keywords:

Photo-induced reactions

Solar energy

Mechanism

Fundamentals

Doping

Graphene

Energy and environmental

Air pollution

Sustainable

Photovoltaic

Hydrogen production

Tutorial review

ABSTRACT

The remarkable achievement by Fujishima and Honda (1972) in the photo-electrochemical water splitting results in the extensive use of TiO₂ nanomaterials for environmental purification and energy storage/conversion applications. Though there are many advantages for the TiO₂ compared to other semiconductor photocatalysts, its band gap of 3.2 eV restrains application to the UV-region of the electromagnetic spectrum ($\lambda \leq 387.5$ nm). As a result, development of visible-light active titanium dioxide is one of the key challenges in the field of semiconductor photocatalysis. In this review, advances in the strategies for the visible light activation, origin of visible-light activity, and electronic structure of various visible-light active TiO₂ photocatalysts are discussed in detail. It has also been shown that if appropriate models are used, the theoretical insights can successfully be employed to develop novel catalysts to enhance the photocatalytic performance in the visible region. Recent developments in theory and experiments in visible-light induced water splitting, degradation of environmental pollutants, water and air purification and antibacterial applications are also reviewed. Various strategies to identify appropriate dopants for improved visible-light absorption and electron–hole separation to enhance the photocatalytic activity are discussed in detail, and a number of recommendations are also presented.

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1. Introduction

Photocatalysis refers to the acceleration of a chemical reaction in the presence of substances called photocatalysts, which can absorb light quanta of appropriate wavelengths depending on the band structure [1–4]. Usually semiconductors including TiO_2 , Fe_2O_3 , WO_3 , ZnO , CeO_2 , CdS , Fe_2O_3 , ZnS , MoO_3 , ZrO_2 , and SnO_2 are selected as photocatalysts due to their narrow band gap and distinct electronic structure (unoccupied conduction band and occupied valence band) [5–24]. In semiconductor photocatalysis, electrons from the valence band of a semiconductor are excited to the conduction band by light of higher energy than the respective band gap, resulting in the formation of $\text{e}^-_{\text{CB}}/\text{h}^+_{\text{VB}}$ pairs (Fig. 1). Conduction band electrons are good reducing agents (+0.5 to –1.5 V vs. NHE) whereas the valence band holes (h^+_{VB}) are strong oxidizing

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