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Photocatalytic performance enhancement of CuO/Cu_2O heterostructures for photodegradation of organic dyes: Effects of CuO morphology

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ABSTRACT

Cuprous oxide is a promising candidate for photocatalysis. But its photocatalytic properties still need to be much improved for applications. Herein, we report a kind of heterostructures (HCs), i.e. CuO/Cu_2O HCs. Different morphologies of CuO, i.e. nanowires, nanotetrahedra and nanospheres, have been controllably prepared on Cu_2O cubes/octahedra by a facile wet chemical method. All the obtained CuO/Cu_2O HCs have significantly improved photocatalytic activity and stability as compared to Cu_2O. Especially, the nanowires CuO/Cu_2O have shown a specific reaction rate ca. 1.6 $\mu mol\ min^{-1}\ g^{-1}$, 260 times as high as pristine Cu_2O and 4 times as Au/Cu_2O. And the nanospheres CuO/Cu_2O have maintained over 95% of photocatalytic activity after 7 cycles. So, distinct morphologies of CuO have resulted in dramatic effects on photocatalytic properties of CuO/Cu_2O HCs.

1. Introduction

Photocatalytic materials have promising applications in environmental remediation. But they have to meet high requirements, i.e. visible light absorption, good charge separation, available reaction sites accessibility and earth-abundance as well as environment benignity [1–4]. Since synthesis of Cu_2O with well-defined morphologies has become controllable, cuprous oxide has attracted great attentions for photocatalysis due to its non-toxicity and earth-abundance [5]. However, photocatalytic activity and stability of Cu_2O still need to be much improved. Aiming for solving the above problems, several strategies have been taken to enhance photocatalytic performances of Cu_2O.

One strategy is to form heterostructures (HCs) by combining cocatalysts like metal oxides, r-GO etc [6,7]. Take a few for examples. Core–shell TiO_2/Cu_2O [8,9], α-Fe_2O_3/Cu_2O [10], sandwich-structured $Cu/Cu_2O/CuO$ as photocathodes [11], and Fe_xO_y nanosheets/Cu_2O cubes [12]. have been reported. Those HCs showed better photocatalytic activities than Cu_2O alone. Very recently, molecular catalyst as a cocatalyst has also been reported [13]. Grätzel reported immobilization of a molecular catalyst on Cu_2O as photocathode for CO_2 reduction, which effectively enhanced photocatalytic reduction efficiency. The other strategy is to assemble cubic and rhombic dodecahedral Cu_2O and the ensembles showed enhanced optoelectronic properties as compared to disordered Cu_2O [14]. Another strategy is plasmonic nano-metals enhancement of Cu_2O photocatalytic activity [15–17]. Au/Cu_2O and Ag/Cu_2O demonstrated superior photocatalytic performances.

Based on energy band theory, CuO and Cu_2O also can form HCs with a type II structure, which facilitates charge separation improving photocatalytic activity. But to prepare such allotropic CuO/Cu_2O HCs has been rarely investigated, to the best of our knowledge. Herein, we report controllable preparation of CuO/Cu_2O HCs and explore the HCs enhance photocatalytic properties. Different morphologies of CuO, i.e. nanowires (NW), tetrahedra (TH), and nanospheres (NS), have been controllably prepared on Cu_2O cubes/octahedra. Enhancements of photocatalytic activity and stability by distinct morphologies of CuO/Cu_2O HCs have been investigated. By comparison, the effect of CuO morphology on photocatalytic activity and stability of CuO/Cu_2O HCs was studied. Based on the above studies, possible photocatalytic mechanism and HCs formation mechanism were hypothesized.

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2. Experimental

2.1. Synthesis of Cu_2O cubes and octahedra

All of the chemical reagents used in this experiment were analytical grade and used without further purification. Cu_2O cubes and octahedra were prepared according to a previously reported method [15]. In a typical procedure, 5 mL of 200 mM NaOH aqueous solution was added dropwisely to 50 mL of 0.01 M $CuCl_2$ aqueous solution at 55 °C. For preparation of octahedral Cu_2O, 1.667 g polyvinylpyrrolidone (PVP, MW = 55,000) was added. After stirring for 0.5 h, 5 mL of 0.6 M ascorbic acid aqueous solution was added dropwisely to the above solution. The mixed solution was kept at 55 °C with vigorous stirring for a certain time (cubes for 5 h, octahedra for 3 h). Finally, the mixed solution gradually became brick-red, indicating formation of Cu_2O. The obtained products were washed with DI water and absolute ethanol several times, and dried in a vacuum oven at 40 °C for 6 h.

2.2. Synthesis of CuO/Cu_2O HCs

2.2.1. NW CuO/Cu_2O

Typical NW CuO/Cu_2O sample was synthesized as follows. A 43.3 mg sample of Cu_2O cubes or octahedra was added to 10 mL of 200 mM NaOH aqueous solution with ultrasonication. And then, the mixed solution was magnetically stirred at room temperature (RT) for 3 h. The as-obtained product was collected by centrifugation, washed with distilled water and absolute ethanol several times. Finally, the product was dried in a vacuum oven at 40 °C for 6 h.

2.2.2. TH CuO/Cu_2O

Typical TH CuO/Cu_2O sample was synthesized as follows. A quantity of 43.3 mg of Cu_2O cubes was added to 5 mL of 0.2 M NaCl or $NaNO_3$ aqueous solution with ultrasonication. And then, the suspension solution was reacted at room temperature for 5 h. After the reaction, the product was collected by centrifugation, washed with DI water and absolute ethanol for several times respectively. Finally, the product was dried in a vacuum oven at 40 °C for 6 h.

2.2.3. NS CuO/Cu_2O

Typical NS CuO/Cu_2O sample was synthesized as follows. A quantity of 43.3 mg of Cu_2O cubes was added to 30 mL of 100 mM EDA aqueous solution with ultrasonication. And then, the suspension solution was sealed into an autoclave vessel and heated at 180 °C for 5 h. After the reaction, the product was collected by centrifugation, washed with DI water and absolute ethanol for several times respectively. Finally, the product was dried in a vacuum oven at 40 °C for 6 h.

2.2.4. Au/Cu_2O

Au/Cu_2O cubes sample was prepared according to our previously reported method [15]. Typical procedure is as follows. 20 mg Cu_2O cubes sample was dispersed into 3 mL distilled water by ultrasonication. And then the suspension was irradiated under a 300 W Xe lamp (PLS-SXE300). Under continuous stirring, 2.0 mL of 0.0625 g/L $HAuCl_4$ aqueous solution was added to the above suspension, and the suspension was further irradiated for 1 h. After that, the products were collected by centrifugation, washed with distilled water and absolute ethanol several times, and dried in a vacuum oven at 40 °C for 6 h.

2.2.5. Control experiment: oxidation of Cu_2O cubes in air

60 mg of Cu_2O cubes was heated up to 400 °C in air at a rate of 5 °C min^{-1} and kept at 400 °C for 1 h before it was allowed to cool down to room temperature.

2.3. Characterizations

The X-ray diffraction(XRD) patterns of the products were measured by using a Bruker D8 Discover diffractometer with Cu-K_α radiation (λ = 0.15406 nm) at a scanning rate of 0.02 deg/s in a 2θ range from 20 to 80°. Scanning electron microscopy (SEM) images were obtained using a Hitachi S4800 scanning electron microscopy. Transmission electron microscopy (TEM) images were recorded on a FEI Tecnai G2 F20 S-TWIN microscope operating at 200 kV. Elemental analysis of the samples was conducted using an energy-dispersive X-ray analysis system attached to the TEM. UV–vis absorption spectra were recorded on a PE Lambda 650 s UV–vis spectrometer. X-ray photoelectron spectra were recorded on a Escalab 250 Xi X-ray photoelectron spectroscopy (XPS).

2.4. Photocatalytic activity and stability tests

Photocatalytic tests were carried out as follows. 15 mg of photocatalyst sample was dispersed into 50 mL of 2×10^{-5} M methyl orange (MO) aqueous solution. The suspended solution was magnetically stirred in the darkness for 0.5 h to reach an adsorption/desorption equilibrium. And then, the mixed solution was irradiated under a 300 W Xe lamp (PLS-SXE300, 150 mw/cm^2) equipped with UV filter (cutoff wavelength λ = 420 nm) from a distance of ca. 15 cm. At a given time interval, 1 mL aliquot of the mixed solution was taken out and centrifuged. And then photocatalytic activity was tested according to absorption spectra of MO, recorded on a UV–vis absorption spectrophotometer (PE Lambda 650s).

3. Results and discussion

3.1. NW CuO/Cu_2O

Fig. 1 shows SEM, TEM images and XRD patterns of NW CuO/Cu_2O made by mild etching of 200 mM NaOH aqueous solution. Morphology and optical property of Cu_2O cubes are shown in Fig. S1. Cu_2O cubes with an edge length ca. 585 nm show good monodispersity. UV–vis absorption spectrum of Cu_2O shows a peak at λ = 484 nm attributed to band absorption and a peak at λ = 550 nm attributed to scattering [18]. In Fig. 1, CuO exhibits nanowire morphology with a diameter of ca. 60 nm and a length of a couple of micrometers. The CuO NWs lie down on the surfaces of Cu_2O cubes. The lattice spacing of 0.232 nm in high resolution TEM (HRTEM) image is assigned to (111) plane of CuO. So, HRTEM image and selective area electron diffraction (SAED) pattern prove the nanowires are CuO and crystalline. Our XPS results in Fig. S2 also show the formation of CuO on surface, in good agreement with HRTEM images. X-ray diffraction (XRD) patterns of the obtained CuO/Cu_2O further demonstrate CuO NWs and Cu_2O in HCs are crystalline. In order to prove feasibility of this method, we did a series of experiments and found surface density of CuO NWs increased with increasing concentration of NaOH aqueous solution, as shown in Fig. S3. CuO morphology changed from nanoflakes to nanowires to flower-like assemblies with the increase of NaOH concentration, clearly indicating NW CuO crystal growth process. FTIR spectra of the samples obtained at different reaction time show intermediate $Cu(OH)_2$ formed and its peak intensity decreased as the reaction time elapsed, as shown in Fig. S4. This is in good agreement with an earlier literature, where Cu_2O reacted with NaOH aqueous solution to form $Cu(OH)_2$ nanowires [19]. Also, this suggests transition from $Cu(OH)_2$ to CuO since $Cu(OH)_2$ easily turns into CuO upon exposed to air and drying.