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Room-temperature broadening and pressure-shift coefficients in the v_2 band of CH₃D–O₂: Measurements and semi-classical calculations

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Abstract

We report measured Lorentz O₂-broadening and O₂-induced pressure-shift coefficients of CH₃D in the v₂ fundamental band. Using a multispectrum fitting technique we have analyzed 11 laboratory absorption spectra recorded at 0.011 cm^{-1} resolution using the McMath-Pierce Fourier transform spectrometer, Kitt Peak, Arizona. Two absorption cells with path lengths of 10.2 and 25 cm were used to record the spectra. The total sample pressures ranged from 0.98 to 339.85 Torr with CH₃D volume mixing ratios of 0.012 in oxygen. We report measurements for O_2 pressure-broadening coefficients of 320 v_2 transitions with quantum numbers as high as J'' = 17 and K = 14, where $K'' = K' \equiv K$ (for a parallel band). The measured O₂-broadening coefficients range from 0.0153 to $0.0645 \text{ cm}^{-1} \text{ atm}^{-1}$ at 296 K. All the measured pressure-shifts are negative. The reported O₂-induced pressure-shift coefficients vary from about -0.0017 to -0.0068 cm⁻¹ atm⁻¹. We have examined the dependence of the measured broadening and shift parameters on the J", and K quantum numbers and also developed empirical expressions to describe the broadening coefficients in terms of m (m = -J'', J'', J'')and J'' + 1 in the ${}^{O}P$ -, ${}^{O}Q$ -, and ${}^{O}R$ -branch, respectively) and K. On average, the empirical expressions reproduce the measured broadening coefficients to within 4.4%. The O₂-broadening and pressure shift coefficients were calculated on the basis of a semiclassical model of interacting linear molecules performed by considering in addition to the electrostatic contributions the atom-atom Lennard-Jones potential. The theoretical results of the broadening coefficients are generally larger than the experimental data. Using for the trajectory model an isotropic Lennard-Jones potential derived from molecular parameters instead of the spherical average of the atom-atom model, a better agreement is obtained with these data, especially for $|m| \leq 12$ values (11.3% for the first calculation and 8.1% for the second calculation). The O₂-pressure shifts whose vibrational contribution are either derived from parameters fitted in the ^{Q}Q -branch of selfinduced shifts of CH_3D or those obtained from pressure shifts induced by Xe in the v_3 band of CH_3D are in reasonable agreement with the scattered experimental data (17.0% for the first calculation and 18.7% for the second calculation). © 2005 Elsevier Inc. All rights reserved.

Keywords: Monodeuterated methane; CH₃D; O₂-broadening; Fourier transform infrared spectroscopy; Spectral lineshape; Semi-classical calculation

1. Introduction

Laboratory spectroscopic studies of the v_2 vibrational band of CH₃D are needed for the correct interpretation of Earth's atmospheric spectra. The v_2 band of CH₃D is the lowest band ($v_0 \approx 2200 \text{ cm}^{-1}$) of a polyad of nine interacting vibrational states. This paper is a continuation of our study of spectroscopic line parameters in this fundamental band [1,2]. Few measurements and calculations of O₂-broadening coefficients for CH₃D have been published and they refer only to the v_3 fundamental [3,4]. To the best

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of our knowledge, no prior measurements of O₂-induced pressure shift coefficients have been made.

Several theoretical efforts concerned with O₂-broadening coefficients of CH₃D have been put forth. Tejwani and Fox [5] calculated N₂-, O₂-, self-, and H₂-broadening coefficients for CH₃D using the Anderson-Tsao-Curnutte (ATC) theory by considering only the electrostatic interaction between the quadrupole moment of the perturber and the octople moment of CH₃D, with an effective value of this octopole moment strongly overestimated (5.6 D $Å^2$, instead of 3.1 D $Å^2$ in this study). They have computed pressure-broadening coefficients for the above broadening gases for a wide range of J and K values as well as the temperature dependence of the pressure-broadening coefficients in the 100-300 K range. The O₂-broadening coefficients of CH_3D in the v_3 band reported in [3] were calculated using the theoretical model developed in [6]. For the intermolecular potential a simple formulation with two adjustable parameters [7] corresponding to dispersion, induction, and repulsive contributions has been used in addition to the electrostatic interactions. Although the calculated results are generally in reasonable agreement with the experimental data, except at high J values where they are underestimated, the two adjustable potential parameters used have probably no physical meaning. Here, we consider a different theoretical approach for the calculation of broadening coefficients that was recently applied to CH₃D self-perturbed [1], CH₃D-H₂ [8], and CH₃D-N₂ [2] without any adjustable parameters. CH₃D is approximated as a linear molecule for its interactions with O₂ and the potential involves the atom-atom Lennard-Jones model [9] in addition to the electrostatic contributions. The theoretical formulation is similar to that described previously [2] and only some features and characteristics of the potential used will be briefly presented.

2. Experimental details

The 11 absorption spectra used in this work were recorded at an unapodized resolution of 0.0056 cm^{-1} using the McMath-Pierce Fourier transform spectrometer (FTS)

Table 1 Experimental conditions for the spectra using in this study

of the National Solar Observatory (NSO) on Kitt Peak. Two absorption cells with path lengths of 10.2 and 25 cm were used in the experiments. The spectra covering the 2048-2318 cm⁻¹ were simultaneously fit in segments of $5-10 \text{ cm}^{-1}$ wavenumber intervals using a multispectrum nonlinear least-squares procedure [10]. This analysis tool allowed us to combine spectra recorded with low pressures of 98% pure CH₃D, self-broadened spectra, and lean mixtures ($\sim 1\%$) of CH₃D in O₂ in a single least-squares solution. The experimental conditions for each spectrum are presented in Table 1. The experimental setup and data reduction methods were described in detail in our earlier study [1]. Residual water vapor lines appeared in all of our spectra as a combination of low pressure absorption due to small amounts (50 mTorr) of residual water vapor present in the vacuum tank enclosing the FTS, and H₂O absorption arising from water vapor in the N2-purged atmospheric paths between the source and the absorption cell and between the absorption cell and the entrance aperture of the interferometer. The wavenumber calibration for the CH₃D line positions was performed with respect to the well-known line centers [11] of residual water lines.

An example of a multispectrum fitted interval is presented in Fig. 1. The top panel shows 11 overlaid spectra of ${}^{Q}R$ -branch transitions with J'' = 11 and K values from 0 to 10. The bottom panel indicates the fit residuals (observed minus calculated for all 11 spectra). The positions of transitions due to contaminants such as residual CO were determined prior to using the multispectrum fitting procedure. During the analysis these lines were included in the least-squares solution using $0.075 \text{ cm}^{-1} \text{ atm}^{-1}$ at 296 K as O₂-broadening and $-0.005 \text{ cm}^{-1} \text{ atm}^{-1}$ as O₂-induced pressure shift coefficients, respectively.

3. Results and discussion

The results of our multispectrum analysis are presented in Table 2. The O₂-broadening coefficients for CH₃D transitions in the v_2 band are given in units of cm⁻¹ atm⁻¹ at 296 K. For each transition we list the line center positions retrieved from the fits (in cm⁻¹) and the transitions

Temperature (K)	Broadening gas	¹² CH ₃ D volume mixing ratio	Path (cm)	Pressure (Torr)
300.20	¹² CH ₃ D	1.00	10.2	0.98
300.00	¹² CH ₃ D	1.00	10.2	2.95
300.90	¹² CH ₃ D	1.00	10.2	17.50
300.70	¹² CH ₃ D	1.00	10.2	79.00
300.20	¹² CH ₃ D	1.00	10.2	110.00
299.80	¹² CH ₃ D	1.00	10.2	152.00
299.80	¹² CH ₃ D	1.00	10.2	303.00
295.90	O_2	~ 0.012	25.0	339.85
295.65	O_2	~ 0.012	25.0	302.50
295.45	O_2	~ 0.012	25.0	202.35
295.45	O_2	~ 0.012	25.0	102.40

Note. 760 Torr = 1 atm = 101.325 kPa.

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