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Standard enthalpy of formation of diphenyl oxide

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Abstract

The energy of combustion of a pure sample diphenyl oxide in crystal and liquid states at $T = 298.15$ K was determined by isoperibolic calorimeter combustion. The enthalpy of fusion, the triple-point temperature and the purity of the sample were estimated by studying the melting process by calorimetric method. The enthalpies of formation in the crystalline, liquid and gaseous states were calculated. All results are compared with the experimental values reported in

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