Diamond & Related Materials 88 (2018) 18–22

Contents lists available at ScienceDirect

Diamond & Related Materials

journal homepage: www.elsevier.com/locate/diamond

High surface area carbon materials derived from corn stalk core as electrode for supercapacitor

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ARTICLE INFO

Keywords:
Activated carbon
Corn stalk core
Electrochemistry
Supercapacitor

ABSTRACT

The activated carbon materials from the corn stalk core raw materials were prepared through the carbonization and activation process and applied as electrode in supercapacitor. The biomass carbon materials activated under different temperatures were tested by cyclic voltammetry, electrochemical impedance spectroscopy and galvanostatic charge-discharge cycling method. The corn stalk core-derived material activated at 700 °C with the highest specific surface area (2349.89 $m^2 g^{-1}$) has exhibited the maximum specific capacitance of 140 $F g^{-1}$. Further detailed characterization and theoretical analysis have demonstrated that the corn stalk core derived activated carbon anode material can not only enhance the capacity of supercapacitor but also realize the comprehensive utilization of corn stalks.

1. Introduction

With the increasing energy consumption, it is urgent to search for sustainable and renewable energies. Considering the environmental problems brought by fossil fuels, the energy conservation and pollution reduction will be the significant factors in social and economic development. As a pollution-free "GREEN" energy storage system, the electrochemical supercapacitors combine both advantage of batteries and conventional capacitors, and have attracted much attention due to their potential applications ranging from mobile devices to electric vehicles [1, 2]. In recent years, researches in developing supercapacitors are underway to provide ideal power source with excellent power density coupled with rapid charge and discharge capabilities and long cycle life [3, 4].

Based on the energy storage mechanism, supercapacitors can be categorized into electrochemical double layer capacitors (EDLC) and faradic pseudocapacitors. The storage mechanism in EDLC is mainly based on the charged separation at electrode and electrolyte interface, while the faradic pseudocapacitors provide capacitor with the reversible faradic reactions occurring at the electrode surface [5, 6]. The most commonly used electrode materials in supercapacitors are conductive polymer, metal oxide and carbon materials, among them, the commercial applications of conductive polymer and metal oxide are limited owing to their small surface area and high cost.

As the earliest electrode applied in supercapacitors, carbon materials have lots of advantages such as low cost, high conductivity, controllable porous structure and high surface area [7–10]. Activated carbon (AC) [11–13], graphite [14, 15] and carbon nanotube [16] are common electrode materials in EDLC. At present activated carbons are widely used for capacitors according to its high porosity and high surface area that favors good charge accumulation at the interface with the electrolyte and therefore high capacitance can be tapped. In view of low cost and environmental protection, biomass wastes are great substitute for coal and petroleum and mainly studied as raw materials to prepare AC materials (ACs) [17, 18]. Huang et al. [19] prepared the activated carbon fibers (ACFs) using wood sawdust as raw materials, the ACFs showed outstanding electrochemical performance. Dechen Liu [20] prepared the rice husk-based activated carbon by KOH activation and the carbon materials possessed a specific surface area as high as 3263 $m^2 g^{-1}$. Corn stalk, as a kind of common biomass wastes, can be used to prepare electrode materials and applied in supercapacitors, which would solve the environment problem of burning stalks and make full use of resource. In addition, the corn stalk core (CSC) with sponge-like shape has abundant natural porous structure, which can be used to obtain ACs with high surface area and porous structure, and thus boost the capacity of supercapacitors when further used as electrode materials. Cao [21] prepared ACs using corn stalk as raw materials and exhibited the high capacitance of 323 $F g^{-1}$ at a current density of 0.1 $A g^{-1}$. However, the complex prepared procedure which introduced the metallic element has increased the cost and limited development. In this work, the ACs derived from CSC were prepared via carbonizing and activating procedure, which was simple and low cost.

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https://doi.org/10.1016/j.diamond.2018.06.018
Received 15 March 2018; Received in revised form 15 June 2018; Accepted 20 June 2018

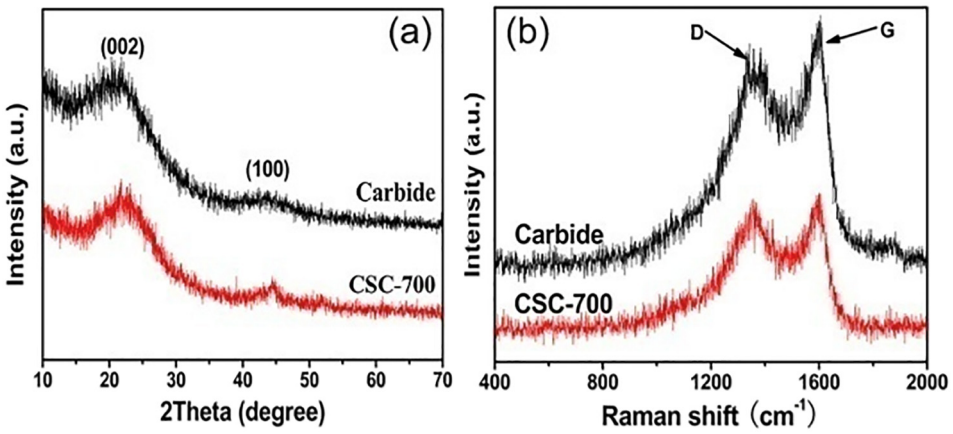

Fig. 1. XRD pattern (a) and Raman spectrum (b) of unactivated carbide and CSC-700.

The ACs activated at different temperatures were applied as electrode in supercapacitor. Moreover, electrochemical measurements were taken to evaluate the performance of various ACs and the optimal activated temperature was selected.

2. Experimental

2.1. Preparation of activated carbon materials

The activated carbon materials (ACs) for this work were synthesized as described follow. Firstly, the corn stalks core (CSC) was cleaned to remove the surface dirt, dried at 60 °C for 24 h and ground into a fine power by pulverizer. Secondly, the dried CSC was carbonized in the tube furnace at 400 °C for 2 h under nitrogen atmosphere with a heating rate of 2.5 °C min^{-1}. Then the obtained carbide was mixed with KOH solution at a mass ratio of 1:3 and dried at 60 °C for 24 h. Later, the dried mixture was heated again at 400 °C under nitrogen atmosphere and held for 30 min before being activated at 500–800 °C for 1 h with a heating rate of 5 °C min^{-1}. Finally, in order to remove the metal-oxide produced in the carbonized process, the obtained activated materials were further purified in 2 mol L^{-1} HCl solution within a water bath at 90 °C for 1 h and washed with deionized water until the solution was neutral, and then dried at 60 °C. The different activated temperature samples were denoted as CSC-500, CSC-600, CSC-700 and CSC-800, respectively.

2.2. Characterization of activated carbon materials

The morphologies and microstructures of ACs samples were observed by scanning electron microscopy (SEM) with a field-emission-scanning electron microscope (JEOL JSM-6700F) and transmission electron microscopy (TEM, JEOL JEM-2100F). Raman spectra were recorded on a Renishaw inVia instrument. X-ray diffraction (XRD) patterns were collected on a Bruker D8 Advance X-ray Diffractometer. The specific surface area and pore size distribution of the carbon materials were measured using nitrogen adsorption–desorption measurements (Micromeritics, ASAP 2420).

2.3. Preparation of electrodes

The working electrode was prepared by mixing the prepared activated materials (ACs), carbon black and polymer binder polyvinylidene fluoride (PVDF) in a 80:10:10 (wt%) ratio in *N*-methyl-2-pyrrolidone. The mixture was uniformly cast on pure nickel foam used as the current collectors and dried under vacuum at 120 °C for 12 h. The coating materials were covered on nickel foam with a surface area of (1 × 1) cm^2. The electrode was pressed under 50 MPa for 5 s. Moreover, in order to be fully soaked, the prepared electrodes were immersed with ethyl alcohol and potassium hydroxide solution for 24 h, respectively. In addition, the platinum electrode and calomel electrode were used as counter electrode and reference electrode, respectively. All the electrodes were dipped into 3 mol L^{-1} potassium hydroxide solution while testing the performance of supercapacitor.

2.4. Electrochemical measurements

In this work, electrochemical measurements were performed on a CHI660E electrochemical workstation. Cyclic voltammetry (CV) measurements were performed with a voltage window of −0.8 V to 0.2 V at different scan rates ranging from 10 mV s^{-1} to 100 mV s^{-1}. Galvanostatic charge-discharge cycling measurements were conducted with a voltage window of −0.8 V to 0.2 V at different current densities ranging from 0.5 A g^{-1} to 5 A g^{-1}. Electrochemical impedance spectroscopy (EIS) measurements were tested with the frequency ranging from 0.01 to 10,000 Hz. The specific capacitance was calculated using the following formula:

$$C_m = \frac{I\Delta t}{\Delta V}$$

In this formula, I is the discharge current density and Δt is the discharge time in the galvanostatic charge-discharge cycling measurements, ΔV is the range of voltage change.

3. Results and discussion

3.1. Characterization of carbon materials

The X-ray diffraction pattern of unactivated carbide and CSC-700 are displayed in Fig. 1(a). Curves possessed no evident sharp diffraction peaks, which indicated the amorphous carbon structure of the carbon materials. Two broad diffraction peaks located at 23° and 44° signified the characteristic peak of graphite (002) lattice plane that attributed to the interconnection and parallel stacking of flake graphite layers, while the (100) lattice suggested the hexagonal honeycomb structure contained in the pyrolytic carbon [22, 23]. Furthermore, the diffraction peak at 44° of CSC-700 was more prominent than that of unactivated carbide, which demonstrated the increasing of graphitization degree after activation.

As shown in Fig. 1(b), both samples exhibited a disorder-induced D-band (~1340 cm^{-1}) and in-plane vibrational G-band (~1590 cm^{-1}) in the Raman spectra, the former ascribed to the defections and disordered carbon in the graphitic structure, and the latter attributed to the vibration of sp2-bonded carbon atoms in a 2D hexagonal lattice. The ratio of intensity between D-band and G-band (ID/IG) reflected the degree of graphitization and the higher ratio represented the lower degree of