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Short communication

Micro-fibrous organic radical electrode to improve the electrochemical properties of organic rechargeable batteries

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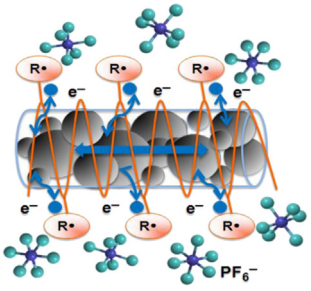
HIGHLIGHTS

- A micro-fibrous organic electrode film (MPE) has been prepared by electrospinning.
- The MPE based lithium cell displays a high discharge capacity of 111 mAh g^{-1} at 1C.
- The MPE has remarkably improved the volumetric capacity as 209 mAh cm^{-3} at 10C.

GRAPHICAL ABSTRACT



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ABSTRACT

A micro-fibrous organic electrode has been prepared by electrospinning process using poly(2,2,6,6-tetramethylpiperidinyloxy-4-ylmethacrylate) (PTMA), poly(vinylidene fluoride-co-hexafluoropropylene) (P(VdF-HFP)) and carbon black powder. The micro-fibrous PTMA electrode (MPE) film is investigated by FT-IR, TGA, SEM, TEM and electrochemical tests. The MPE shows excellent rate-capability and cycle performance. The MPE based lithium cell displays a high discharge capacity of 111 mAh g^{-1} at 1C. Even at a high current density of 50C, the MPE cell presents a discharge capacity of 109 mAh g^{-1}. Moreover, the MPE have remarkably improved the volumetric capacity (209 mAh cm^{-3} at 10C) compared with a nano-fibrous PTMA electrode. The excellent electrochemical performance is ascribed to the microstructure that promotes fast ion transport through short diffusion pathways and at the same time facilitates electron transport.

1. Introduction

Organic rechargeable batteries (ORBs) have been on the rise in recent years driven by their potential applications in flexible electronic devices. Beside the flexibility, the ORB is beneficial in terms of easy recycling/disposal, being environmentally friendly, utilizing unlimited material sources, being lightweight, and offering a high power density. Also, the energy density could be dramatically increased by design of the molecular structure. An early attempt to develop ORB used conducting polymers such as poly(acetylene) (PAC), polypyrrole (PPY) and polyaniline (PANI), but these batteries generally showed low capacity and poor cycleability [1,2]. Recently, carbonyl, sulfuric, organic crystal, and 2,2,6,6-tetramethylpiperi dinyloxy (TEMPO)-based organic materials have attracted attention as cathode for ORB [3–15].

Among the above materials, poly(2,2,6,6-tetramethylpiper idinyloxy-4-ylmethacrylate) (PTMA) containing the TEMPO radical is known to show redox behavior at 3.6 V vs. Li/Li+, at low

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rates. It also provides high power and rate capability with a theoretical cathode capacity of 111 mAh g^{-1}, due to a fast electron transfer rate [12]. However, a large amount carbon conductor has been necessary in the electrode, because PTMA is an insulator. Moreover, a poor contact with carbon conductor has been observed at high amounts of active material; with PTMA being dissolved into organic electrolytes and inducing self-discharge [16–18]. To improve the rate capability at high contents of active material graphene and CNT composite PTMA electrodes have been prepared, but cannot prevent the dissolution of active material [19,20]. Some strategies, such as using crosslinking, brush framework, mesocellular carbon, and ionic liquid etc [18,21–23], have been reported in order to dissolve the dissolution problem of PTMA.

We have also proposed a new strategy, a nano-fibrous PTMA film fabricated by electrospinning, to overcome the above-mentioned negative features of PTMA [24]. Although the nanofibrous morphology film has extraordinary rate capability, excellent cycleability, enhanced flexibility, and prevents dissolution, the nanostructure decrease the tap density and result in low volumetric energy density that needs to be considered in fabricating conventional batteries in industry.

Therefore, in this study, a micro-sized PTMA fibrous electrode film (hereafter MPE) is prepared to increase the volumetric capacity at high PTMA loadings ($\sim$9.0 mg cm^{-2}). The MPE shows improved volumetric capacity with excellent rate capability.

2. Experimental

PTMA was synthesized by the radical polymerization method as described in a previous study [16,18]. 2,2,6,6-Tetramethylpiperidine methacrylate monomer was polymerized by first using 2,2′-azobisisobutyronitrile radical initiator and then oxidizing with H_2O_2 in the presence of a $NaWO_4$ catalyst to obtain PTMA. 60 wt.% PTMA, 10 wt.% poly(vinylidene fluoride-co-hexafluoropropylene) (P(VdF-HFP)), and 30 wt.% carbon black powder were added to acetone and N-methylpyrrolidinone (1:1, w/w) solution. Electrospinning was performed by applying a voltage of 20 kV at room temperature. The electrospun micro-sized PTMA fibers were collected on an aluminum current collector to form an electrode film and vacuum dried at 80 °C for 12 h before further use [24]. An MPE based ORB was fabricated by stacking a lithium metal (300 μm thickness, Cyprus Foote Mineral Co.) anode with a Celgard 2200 separator film, and 1 M $LiPF_6$ in ethylene carbonate (EC)/dimethyl carbonate (DMC) (1:1 v/v) (Supplied from Samsung Co.) as the electrolyte. Electrochemical performance tests were carried out using an automatic galvanostatic charge–discharge unit, WBCS3000 battery cycler, between 3.0 and 4.0 V at room temperature, at various current densities. The size distribution of MPE was observed by transmission electron microscopy (TEM) performed on a JEM-2010 JEOL and scanning electron microscopy (SEM) using Philips XL30 S FEG. FT-IR absorption spectra were recorded with a Fourier transform interferometer (VERTEX 80v, Bruker Optics). The thermal stability was analyzed by thermogravimetric analysis (TGA) (SDT Q600 TA) in a nitrogen atmosphere at a heating rate of 10 °C min^{-1} from 20 °C to 800 °C.

3. Results and discussion

FT-IR spectra of PTMA, PVdF-HFP, and MPE are presented in Fig. 1. The main characteristic peaks of PVdF-HFP were assigned as follows, 1400, 1200 and 480 cm^{-1} attribute to CF_2 bending, CF_2 stretching, and CF_2 wagging [25]. The spectrum of PTMA shows absorption at 1728 and 1360 cm^{-1}, which corresponds to C=O and NO•. The spectrum of MPE has the characteristic peaks of both PTMA and PVdF-HFP, and broad peaks are displayed by addition of carbon. The peaks of MPE are almost the same with the precursors because of the absence of interactions between precursors.

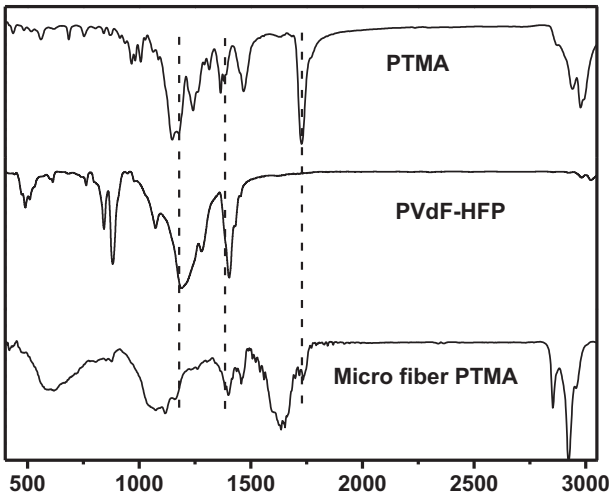

Fig. 1. FT-IR spectrum of PTMA, PVdF-HFP and micro-fiber PTMA electrode film.

The good thermal stability of MPE is revealed by TGA analysis of Fig. 2. An initial decomposition of PTMA occurs at 270 °C and 52 wt.% of PTMA in MPE is disappeared at 430 °C [26]. The initial decomposition of PVdF-HFP occurs at 430 °C and the decomposition of PVdF-HFP is completed at 500 °C with 8 wt.% loss. The dried carbon of 30 wt.% is maintained to 500 °C, and $\sim$10 wt.% ash from PTMA and PVdF-HFP is formed at 500 °C.

In Fig. 3a SEM image of MPE shows an interconnected network of long and continuous micro fibres with a diameter distribution from 1 μm to 5 μm. The high magnification of the inserted figure reveals the thick portion at the fiber caused by aggregation of carbon powder. The TEM image of Fig. 3b exhibits that the carbon black phase of the fiber is located mainly in the center of the organic fibers, although some aggregation of carbon black particles occur. The polymer layer on carbon black has a diameter distribution from 0.5 μm to 1 μm.

A simplified schematic of the micro organic fiber is given in Fig. 4a. The fully interconnected micron-sized pores in the film are filled with electrolyte. The nitroxyl radical of TEMPO is oxidized to form a cation and joins a PF_6^- of the electrolyte to form an

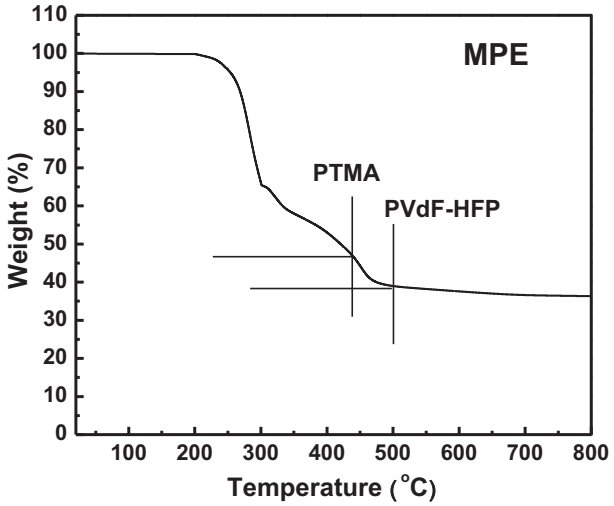

Fig. 2. TGA curve of micro-fibrous PTMA electrode film (MPE).