



Short communication

Cycle versus voltage hold – Which is the better stability test for electrochemical double layer capacitors?

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HIGHLIGHTS

- ▶ Stability tests for double layer capacitors by cycling and voltage hold were compared.
- ▶ The voltage hold test turns out to be more demanding than the cycle test.
- ▶ The same capacitance loss was observed after much shorter test time for the voltage hold test.
- ▶ The cycle test overestimated the voltage limit for stability by about 0.3 V.

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ABSTRACT

The stability of electrochemical double layer capacitors (EDLC) utilizing an ionic liquid and tetraethylammonium tetrafluoroborate (TEABF₄) in acetonitrile electrolyte was tested by cycling and by constant voltage hold tests. It turned out that possible degradation of the EDLC is visible in the voltage hold test after a time period which is typically much shorter if compared to the cycle tests. A constant capacitance during several thousands of cycles does not necessarily indicate stability. For capacitor cells using a carbon black as active material and an ionic liquid as electrolyte constant capacitance could be demonstrated over 12,000 cycles between 0 V and 3.75 V. Clear ageing was observed, however, during the voltage hold test at 3.75 V after 300 h. For a maximum voltage of 3.5 V both the cycle and the voltage hold test indicated stable performance over 12,000 cycles or 500 h, respectively. For capacitors using an activated carbon as active material and standard electrolyte TEABF₄ in acetonitrile both tests clearly demonstrated degradation for a maximum cell voltage of 3.5 V. The evolution of capacitance loss observed for the voltage hold test at 3.25 V is in good agreement with the degradation observed at 3.5 V for the voltage hold test.

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1. Introduction

Electrochemical double layer capacitors (EDLC) are known for their high power density and cycle stability. Commercial products provide a cycle life of one million or above at nominal conditions, which correspond to cycling between 0 V and the nominal cell voltage at room temperature. The end of life criterion is 20% or 30% loss in capacitance or doubling of the internal resistance [1].

In most of the scientific literature the stability of a newly developed EDLC material or EDLC system is demonstrated via cycling and typically several thousands of cycles are presented without significant capacitance loss [2,3]. However, apart from the fact that accumulation of several thousand cycles is already rather

time consuming, the question arises, if several thousand cycles are really sufficient to proof the life time stability of a device over several hundred thousand cycles or even one million?

An alternative to the cycling test is the constant load or float test performed by holding the capacitor at the nominal cell voltage and determining the capacitance as a function of time by occasional charge/discharge cycles [4,5]. Like cycle tests, which may reflect the duty cycle for some applications of EDLC, voltage hold experiments directly reflect conditions for those applications where the capacitor is an emergency or backup device being permanently charged and waiting for the emergency event to occur.

The significant difference between these two test procedures and the respective cycle-rating and float-rating of a capacitor was already pointed out by M. Okamura et al. [6], who stated that during cycling only a “derated” voltage is applied to the capacitor increasing the measured life time. In the present communication we will compare both test procedures and discuss advantages and

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disadvantages of one or the other. Although we assume that our conclusions drawn from tests performed with laboratory scale cells also hold for commercial cells it cannot be excluded that certain differences will occur.

2. Experimental

2.1. Materials

For preparation of the electrodes two high surface area carbons (YP17 (Kuraray Chemicals Japan) and an experimental carbon black (Timcal, Switzerland)) were used for standard electrolyte and ionic liquid, respectively. Both carbons were used as received. Electrodes were prepared by mixing 90 wt. % of carbon material together with 10 wt. % of Polytetrafluoroethylene (PTFE) binder (60% dispersion in water, Sigma Aldrich). The two components were dispersed in a mixture of isopropanol:water 1:1, heated up to 100 °C and constantly stirred until a dough like mass was achieved. Afterwards self-standing electrodes were prepared out of the two different carbon materials. The thickness of the YP17 electrodes is in the range of 200 µm, whereas the thickness of the carbon black electrodes was increased to approximately 500 µm to improve the mechanical stability. The electrode sheets were dried for at least 24 h at 120 °C and a vacuum of 20 mbar before they were assembled in capacitor cells.

Prior to use the ionic liquid [EMIM][BF₄] (purity > 98%, Merck KGaA Darmstadt) was dried in a Schlenck tube at 100 °C and a vacuum of 2×10^{-3} mbar for at least 8 h. The standard electrolytes, i.e. 1 M tetraethylammonium tetrafluoroborate (TEABF₄) in acetonitrile (AN) (Honeywell speciality chemicals Seelze, Germany) was used as received. Karl Fischer titration (KF Coulometer 684, Metrohm, Switzerland) showed for all electrolytes water contents below 20 ppm.

2.2. Electrochemistry

The assembly and filling of the electrochemical cells were described before [4]. In brief, the activated carbon/carbon black electrodes (diameter 12 mm) are sandwiched between two spring loaded (50 Ncm²) titanium pistons. A carbon coated aluminium layer (Gaia Akkumulatorenwerke, Germany) was put in between titanium and electrode for better contact. The two electrodes were separated using a 30 µm thick paper separator (Maxwell Technologies, Switzerland) or a glass fibre separator (EUJ116 from Hollingsworth & Vose, UK) for standard electrolyte and ionic liquid, respectively. Before each test the cells were dried at 120 °C under vacuum for at least 12 h. The cells were vacuum filled using a syringe connected with a tube to the cell body inside a glovebox with O₂ and H₂O contents below 1 ppm (MBraun, Germany).

For the cycling tests a specific current was applied for cycling the full cell between the critical upper voltage and 0 V. In the case of the ionic liquid/carbon black cells a specific current of 0.75 Ag⁻¹ was applied. The cells with standard electrolyte/YP17 activated carbon were cycled with 1 Ag⁻¹. The specific current refers to the mass of both electrodes. The respective capacitance was evaluated from the current and time for discharge, typically each 300 cycles. The available cell voltage was determined considering the IR drop (after approx. 3–4 s) of the cell.

The voltage hold test procedure has been described in detail elsewhere [4]. In brief, the critical voltage was applied to the cell and by occasional cycling the cell between 2.5 V and 0 V with a constant current of 0.28 Ag⁻¹ the respective capacitance was determined. Three cycles were applied every 10 h during the voltage hold procedure. An upper voltage of 2.5 V was considered to be harmless for the cell.

Cycle tests with standard electrolyte/YP17 electrodes were performed with a programmable potentiostat (VMP 3 workstation, Biologic France). The cycle and voltage hold tests with ionic liquid were performed using a CCCC Battery cycler (Astrol Technologies, Switzerland). All tests were performed at room temperature without temperature control.

3. Results and discussion

In order to increase energy and power density of EDLCs the cell voltage should be increased above the typical 2.5–2.8 V of currently available EDLC utilizing activated carbon and TEABF₄ in acetonitrile. Novel electrolytes such as ionic liquids [7,8] or novel carbons such as carbon nanotubes [9,10] or graphene [11,12] or combinations of such [13,14] were suggested to allow for an increased cell voltage without loss in power and stability. The preservation of stability is usually demonstrated by cycling experiments over several tens of thousand cycles between the critical cell voltage and zero voltage. Cycling experiments are relatively easy to set up, but are rather time consuming; especially when considering that 500,000 to 1,000,000 cycles are commonly required for commercial products before a capacitance loss of 30% is observed. Demonstrated 20,000 cycles, therefore, just cover 2–4% of the aimed cycle life.

For example, for the asymmetric EDLC system utilizing pyrrolidinium based ionic liquids as electrolytes at temperatures of 60 °C a potential window of 4.0 V was demonstrated by Mastragostino [15] by cycling. About 20,000 cycles at a rate of 1 Ag⁻¹ were obtained before 20% of the initial capacitance was lost.

Charge/discharge cycle experiments are typically used for batteries to demonstrate stability. Batteries are usually prone to degradation at the upper as well as at the lower voltage limit. Even at intermediate voltages batteries may suffer from charge transfer reactions occurring in the bulk of the electrode materials. In contrast, EDLCs are accepted to be very stable at voltages below the nominal cell voltage and can be discharged to zero without harm.

As a consequence, during cycling experiments between the fully charged state and the discharged state the capacitor is most of the time exposed to voltages, which are known not to affect the stability. In order to avoid this waste of cycling time we suggest that the alternative constant hold test (or constant voltage hold test or float test) is much more efficient for determination of the EDLCs stability limits. This situation is sketched in Fig. 1 where for both tests voltage and current are plotted. The solid (blue) line represents the cycle test while the dotted (red) line shows current (bottom) and voltage (top) versus time for the hold test. The grey areas (Fig. 1, top) clearly show that the time the capacitor is exposed to voltages above an assumed stability limit is small for the cycle test. For the hold test, the capacitor is constantly exposed to voltages above the assumed stability limit. Another fundamental difference between the two tests reveals the current developing (Fig. 1, bottom). The cycling test is performed at an arbitrarily set and constant current. During the voltage hold test the recorded current corresponds to the leakage current, which decreases constantly during the test period [16].

In order to compare both ageing test procedures we measured the capacitance decrease for a capacitor cell utilizing [EMIM][BF₄] and carbon black electrodes. The upper potential limits were set to 3.75 V and 3.5 V. The cycling test was performed by cycling between 0 V and the respective upper cell voltage at a specific current of 0.75 Ag⁻¹. With this specific current one complete cycle takes 300 s. In parallel voltage hold tests at 3.75 V and 3.5 V were performed according to the procedure described in the experimental part and in the publication by Ruch et al. [4].

The results of the cycling tests are reproduced in Fig. 2a and b for the ionic liquid [EMIM][BF₄].

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