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Solid-liquid phase equilibrium and thermodynamic functions: 2-Nitrobenzaldehyde in different neat solvents and ternary system of 2-nitrobenzaldehyde + 4-nitrobenzaldehyde + ethyl acetate

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ABSTRACT

Equilibrium solubility data were determined for binary systems of 2-nitrobenzaldehyde in ethanol, isopropanol, *n*-butanol, *n*-butyl acetate, acetone, ethyl acetate, toluene, acetic acid, cyclohexane, *n*-propanol and for ternary system of 2-nitrobenzaldehyde + 4-nitrobenzaldehyde + ethyl acetate by the static method under atmospheric pressure of 101.1 kPa. The binary systems were measured at $T = (278.15, 280.65, 283.15, 285.65, 288.15, 290.65, 293.15, 295.65 \text{ and } 298.15) \text{ K}$, and the ternary system were determined at the three temperatures of 278.15 K, 288.15 K and 298.15 K. The solubility of 2-nitrobenzaldehyde in neat solvents increased with increasing temperature, and obeyed the following order

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